

This is a **tentative** outline of the sections of the textbook that should be the main emphasis of each day's lecture. Some topics may carry over into the next lecture and, if time permits, we may begin some topics prior to the day mentioned. This outline may change as the semester progresses.

Lecture	Date	Main Topics	Homework and Due Dates
1	M: 8/23	Course outline/Syllabus 1.1 Soda Pop Fizz	
2	W: 8/25	1.2 Chemicals Compose Ordinary Things 1.3 Atoms and Molecules 1.4 The Scientific Method	
3	F: 8/27	2.1 Global Temperature 2.2 Scientific Notation	
4	M: 8/30	2.3 Significant Figures 2.4 Significant Figures in Calculations	
5	W: 9/1	2.5 Basic Units of Measurements 2.6 Problem Solving and Unit Conversions	
6	F: 9/3	2.7 Multistep Problems 2.8 Units Raised to a Power 2.9 Density	Quiz 1 Posted
	M: 9/6	No Classes: Labor Day	
7	W: 9/8	Review Quiz 1	Quiz 1 scantron must be turned in at the beginning of class
8	F: 9/10	Exam 1	

9	M: 9/13	3.1 In Your Room 3.2 What is Matter 3.3 & 3.4 Classifying Matter	
10	W: 9/15	3.5 Differences in Matter 3.6 Changes in Matter 3.7 Conservation of Mass	
11	F: 9/17	3.8 Energy 3.9 Physical and Chemical Changes 3.10 Temperature	
12	M: 9/20	3.11 Heat Capacity 3.10 Calculations with Energy and Heat Capacity	
13	W: 9/22	4.1 Experiencing Atoms 4.2 Atomic Theory 4.3 The Nuclear Atom 4.4 Subatomic Particles	
14	F: 9/24	4.5 Elements 4.6 Periodic Table	Quiz 2 Posted
15	M: 9/27	4.7 Ions 4.8 Isotopes	
16	W: 9/29	Review Quiz 2	Quiz 2 scantron must be turned in at the beginning of class
17	F: 10/1	Exam 2	
18	M: 10/4	5.1 Sugar and Salt 5.2 Constant Composition 5.3 Chemical Formulas	
19	W: 10/6	5.4 Elements and Compounds 5.5 Writing Formulas	
20	F: 10/8	5.6 Nomenclature 5.7 Naming Ionic Compounds	

		5.8 Naming Molecular Compounds	
21	M: 10/11	5.9 Naming Acids 5.10 Nomenclature Summary 5.11 Formula Mass	
22	W: 10/13	6.1 How Much Sodium 6.2 The Pound 6.3 Counting Atoms by the Gram 6.4 Counting Molecules by the Gram	
23	F: 10/15	6.5 Chemical Formulas and Conversion Factors 6.6 & 6.7 Mass Percent Composition	
24	M: 10/18	6.8 Empirical Formulas 6.9 Molecular Formulas	
25	W: 10/20	7.1 Grade School 7.2 Evidence of a Chemical Reaction 7.3 Chemical Equations	
26	F: 10/22	7.4 Balance Chemical Equations 7.5 Aqueous Solutions and Solubility 7.6 Precipitation Reactions	Quiz 3 Posted
27	M: 10/25	7.7 Writing Chemical Equations 7.8 Acid-Base Reactions 7.9 Oxidation-Reduction Reactions 7.10 Classifying Chemical Reactions	
28	W: 10/27	Review Quiz 3	Quiz 3 scantron must be turned in at the beginning of class
29	F: 10/29	Exam 3	
30	M: 11/1	8.1 Climate Change 8.2 Making Pancakes 8.3 Mole to Mole Conversions 8.4 Mass to Mass Conversions	

31	W: 11/3	8.5 Limiting Reactants 8.6 Theoretical and Percent Yield	
32	F: 11/5	8.7 Enthalpy	
33	M: 11/8	9.1 Models of the Atoms 9.2 Light 9.3 Electromagnetic Spectrum	
34	W: 11/10	9.4 Bohr Model 9.5 Quantum-Mechanical Model 9.6 Electron Configurations	
35	F: 11/12	9.7 Electron Configurations using Periodic Table 9.8 Quantum-Mechanical Model Uses 9.9 Periodic Trends	
36	M: 11/15	10.1 Bonding Models 10.2 Valence Electrons 10.3 Lewis Structures (Ionic Compounds)	
37	W: 11/17	10.4 Lewis Structures (Covalent Compounds) 10.5 Writing Lewis Structures	
38	F: 11/19	10.6 Resonance 10.7 Predicting Shapes of Molecules	Quiz 4 Posted
39	M: 11/29	10.8 Electronegativity and Polarity 10.5 Writing Lewis Structures	
40	W: 12/1	Review Quiz 4	Quiz 4 scantron must be turned in at the beginning of class
41	F: 12/3	Exam 4	
42	M: 12/6	11.1 Extra Long Straws 11.2 Kinetic Molecular Theory 11.3 Pressure	

43	W: 12/8	11.4 Boyle's Law 11.5 Charles's Law 11.6 Combined Gas Law	
44	F: 12/10	11.7 Avogadro's Law 11.8 Ideal Gas Law 11.9 Mixture of Gases 11.10 Gases in Chemical Reactions	
	F: 12/17	Final Exam: 10:15 AM - 12:15 PM	